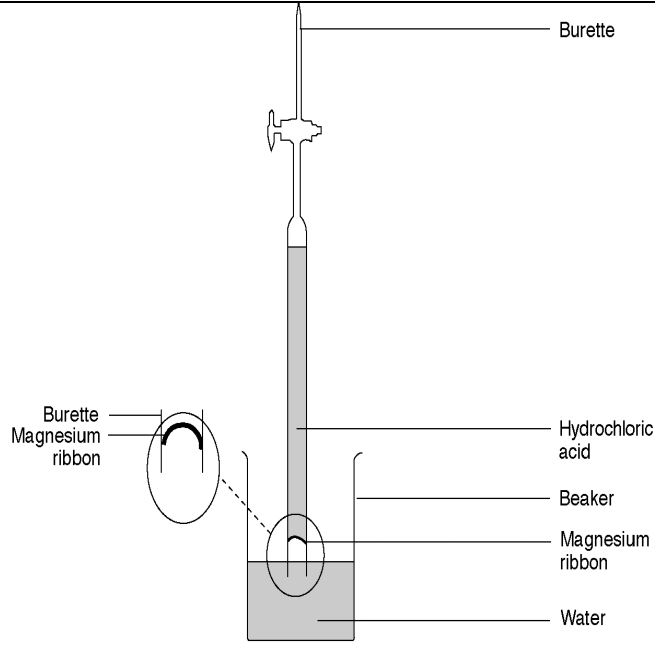


Introduction

One mole of any gas occupies the same volume when measured under the same conditions of temperature and pressure. In this experiment, the volume of one mole of hydrogen is calculated at room temperature and pressure.

	<p>What to do</p> <ol style="list-style-type: none"> 1. Clean a piece of magnesium ribbon about 3.5 cm long and weigh accurately. (This should weigh between 0.02 and 0.04 g; if not adjust the amount used.) 2. Measure 25 cm³ of dilute hydrochloric acid (Irritant) into the burette. Carefully add 25 cm³ of water on top of this. 3. Push the magnesium into the end of the burette so it will stay in position with its own tension. 4. Add 50 cm³ of water to a 250 cm³ beaker. 5. Quickly invert the burette into the water. If this is done quickly and carefully very little is lost. It is important that the liquid level in the burette starts on the graduated scale. If it is not on the scale; momentarily open the tap, this allows the level to drop). Clamp the burette vertically.
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6. Take the burette reading (care: it is upside down!)
7. Observe the magnesium react as the acid diffuses downwards, wait until all the magnesium has reacted.
8. Note the new volume on the burette (care: it is upside down).
9. Record your results.

What to record

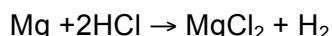
The mass of magnesium used and the volume of hydrogen produced.

Safety

Wear eye protection.

Questions

The equation for the reaction is



The equation tells us that 1 mole of magnesium will produce 1 mole of hydrogen

1. How many moles of magnesium were reacted?
2. How many moles of hydrogen were produced?
3. Calculate the volume of 1 mole of hydrogen.
4. Write a conclusion.
5. How can you make the experiment more valid?